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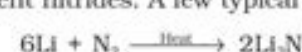
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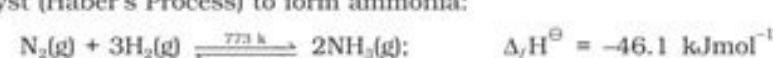
Properties

Dinitrogen is a colourless, odourless, tasteless and non-toxic gas. Nitrogen atom has two stable isotopes: ^{14}N and ^{15}N . It has a very low solubility in water (23.2 cm³ per litre of water at 273 K and 1 bar pressure) and low freezing and boiling points (Table 7.1).

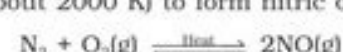
Dinitrogen is rather inert at room temperature because of the high bond enthalpy of $\text{N}=\text{N}$ bond. Reactivity, however, increases rapidly with rise in temperature. At higher temperatures, it directly combines with some metals to form predominantly ionic nitrides and with non-metals, covalent nitrides. A few typical reactions are:



It combines with hydrogen at about 773 K in the presence of a catalyst (Haber's Process) to form ammonia:



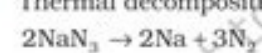
Dinitrogen combines with dioxygen only at very high temperature (at about 2000 K) to form nitric oxide, NO.



Uses The main use of dinitrogen is in the manufacture of ammonia and other industrial chemicals containing nitrogen, (e.g., calcium cyanamide). It also finds use where an inert atmosphere is required (e.g., in iron and steel industry, inert diluent for reactive chemicals). Liquid dinitrogen is used as a refrigerant to preserve biological materials, food items and in cryosurgery.

Example 7.3 Write the reaction of thermal decomposition of sodium azide.

Solution Thermal decomposition of sodium azide gives dinitrogen gas.



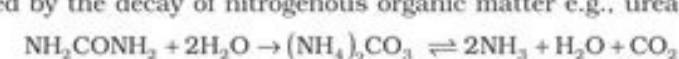
Intext Question

7.3 Why is N_2 less reactive at room temperature?

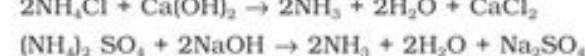
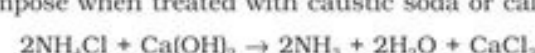
7.3 Ammonia

Preparation

Ammonia is present in small quantities in air and soil where it is formed by the decay of nitrogenous organic matter e.g., urea.



On a small scale ammonia is obtained from ammonium salts which decompose when treated with caustic soda or calcium hydroxide.



178 The p-Block Elements

2020-21

Intext Questions

- 1.9 Give the significance of a 'lattice point'.
- 1.10 Name the parameters that characterise a unit cell.
- 1.11 Distinguish between
 - (i) Hexagonal and monoclinic unit cells
 - (ii) Face-centred and end-centred unit cells.
- 1.12 Explain how much portion of an atom located at (i) corner and (ii) body-centre of a cubic unit cell is part of its neighbouring unit cell.

1.6 Close Packed Structures

In solids, the constituent particles are close-packed, leaving the minimum vacant space. Let us consider the constituent particles as identical hard spheres and build up the three-dimensional structure in three steps.

(a) Close Packing in One Dimension

There is only one way of arranging spheres in a one-dimensional close packed structure, that is to arrange them in a row and touching each other (Fig. 1.17).



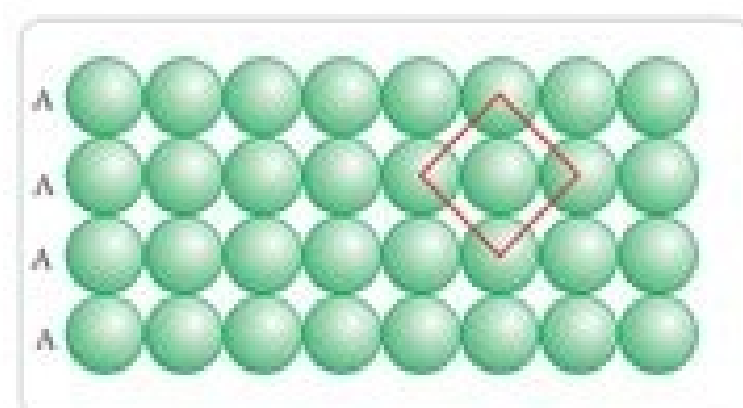
Fig. 1.17: Close packing of spheres in one dimension

In this arrangement, each sphere is in contact with two of its neighbours. The number of nearest neighbours of a particle is called its **coordination number**. Thus, in one dimensional close packed arrangement, the coordination number is 2.

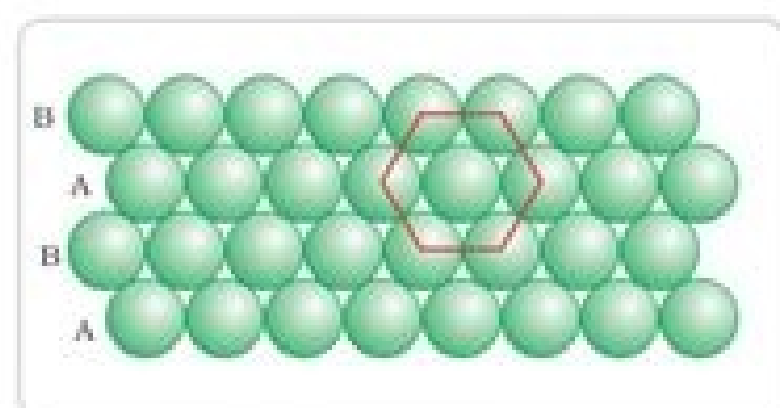
(b) Close Packing in Two Dimensions

Two dimensional close packed structure can be generated by stacking (placing) the rows of close packed spheres. This can be done in two different ways.

- (i) The second row may be placed in contact with the first one such that the spheres of the second row are exactly above those of the first row. The spheres of the two rows are aligned horizontally as well as vertically. If we call the first row as 'A' type row, the second row being exactly the same as the first one, is also of 'A' type. Similarly, we may place more rows to obtain AAA type of arrangement as shown in Fig. 1.18 (a).

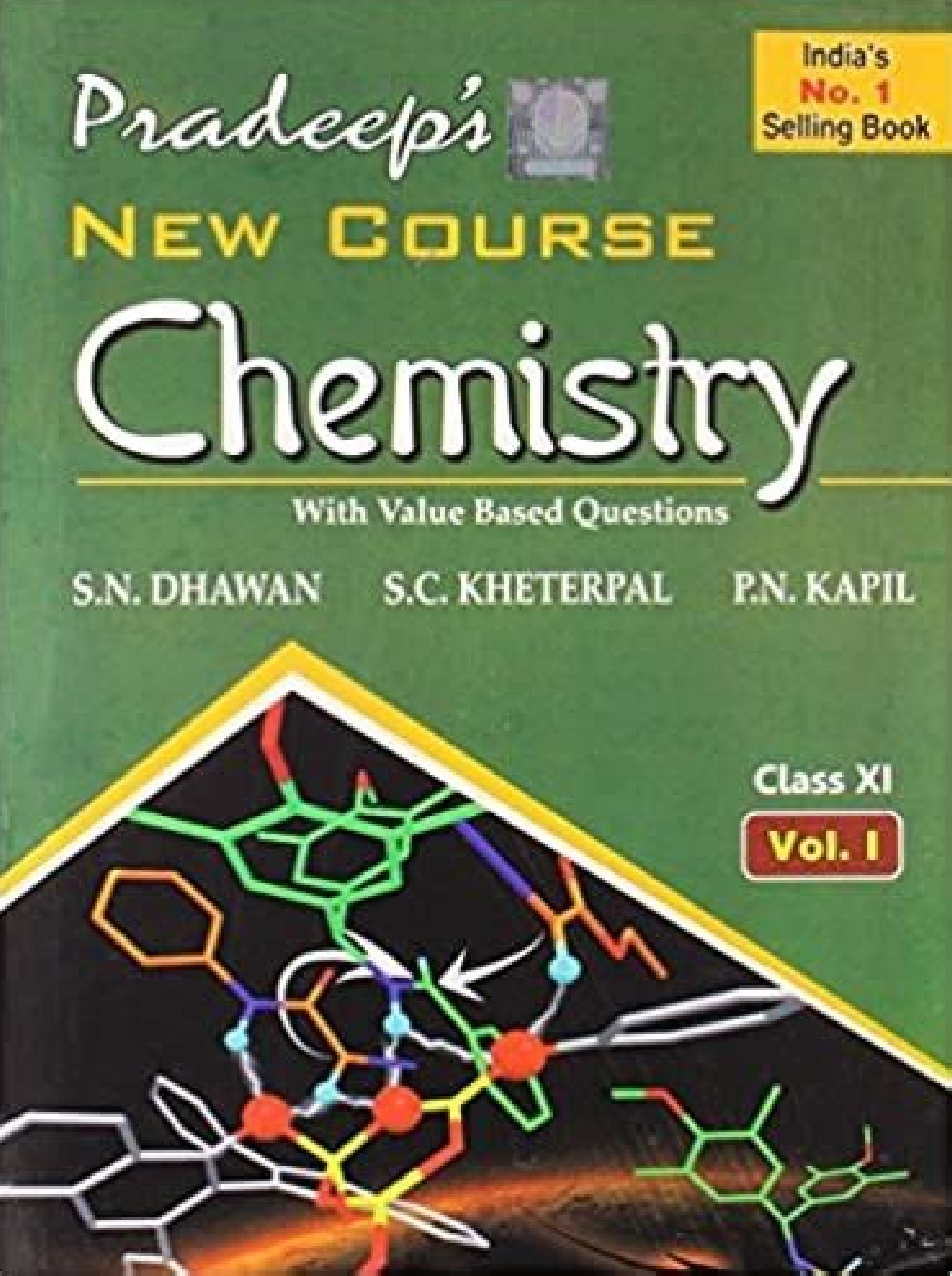


(a)



(b)

Fig. 1.18: (a) Square close packing (b) hexagonal close packing of spheres in two dimensions



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